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Solution:  $n_1 V_1 = n_2 V_2$   
 $0.50 \text{ mmol mL} \times 25.0 \text{ mL} = n_2 \times 120 \text{ mL}$   
 $n_2 = \frac{0.50 \text{ mmol mL} \times 25.0 \text{ mL}}{120 \text{ mL}} = 10.4 \text{ mmol}$   
2. (a) Given:  $V_{\text{HCl(aq)}} = 25.00 \text{ mL}$ ;  $[\text{HCl(aq)}] = 0.350 \text{ mol/L}$ ;  $V_{\text{NaOH(aq)}} = 5.0 \text{ mL}$ ;  $[\text{NaOH(aq)}] = 0.500 \text{ mol/L}$ ; Required:  $[\text{H}^+(\text{aq})]$ , pH  
Analysis: Before titration, amount of  $\text{H}^+$  is:  $0.350 \text{ mmol mL} \times 25.0 \text{ mL} = 8.75 \text{ mmol}$

~~Section 8.7: Acid – Base Titration Tutorial 1 Practice, page 547~~

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Identify any errors in the structure by drawing them. Rename the  
structure correctly. 2-ethyl-2,4,4-trimethylpentane What Is  
Required? You are to identify the errors in the name by drawing the  
structure and renaming it correctly. What Is Given? You are given  
an incorrectly named alkane.

~~Unit 1 Organic Chemistry — Mr. Arthur's Science Page~~  
= 0.12 mol/L Statement: The equilibrium concentration of  
dinitrogen tetroxide gas is 0.12 mol/L. 2. Given: initial quantity of  
NOCl(g) = 3.00 mol; volume = 3.00 L; [NO(g)] initial = 0.00  
mol/L; [Cl<sub>2</sub>(g)] initial = 0.00 mol/L; [NO(g)] equilibrium = 0.043  
mol/L Required: [NOCl(g)] equilibrium; [Cl<sub>2</sub>(g)] equilibrium  
Solution: Step 1.

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~~Section 7.1: Equilibrium Systems Mini Investigation: The ...~~

Solution: bonds broken 3 mol 1 mol CH CCl 3 mol HnD DD --

= x = x + x = 413 kJ mol x + 1 mol 339 kJ mol x H

= 1578 kJ Statement: The energy required to separate 1 mol of chloromethane into free atoms is 1578 kJ. 2. Solution: Step 1: Write the balanced chemical equation for the combustion of 1 mol of ethene gas. C 2H 4(g) + 3 O 2(g) ...

~~Section 5.3: Bond Energies Tutorial 1 Practice, page 312~~

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Solution:  $q = Ne = (1.2 \times 10^8)(1.602 \times 10^{-19} \text{ C})$   $q = 1.92 \times 10^{11} \text{ C}$  (one extra digit carried)  $F E = kq_1 q_2 / r^2 = 8.99 \times 10^9 \text{ N} \cdot \text{m}^2 / \text{C}^2 \cdot \# \text{ } \$ \% \& ' ((1.92 \times 10^{11} \text{ C})^2 / (1.10 \text{ m})^2)$   $F E = 2.7 \times 10^{12} \text{ N}$

Statement: The force of repulsion between the two plastic spheres is

$2.7 \times 10^{12} \text{ N}$ . 2. Given:  $m = 2.48 \times 10^{-15} \text{ kg}$ ;  $d = 1.7 \text{ cm}$

$= 0.017 \text{ m}$ ;  $V_b = 260 \text{ V}$ ;  $e = 1.602 \times 10^{-19} \text{ C}$ ;  $g = 9.8 \text{ m/s}^2$

Required:  $q$ ;  $N$

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