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Periodic Table 04 || Atomic Radius ,Ionic Radius its
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The metallic character of the elements decreases as you move across a Period on the Periodic table, from left to right, and it increases as you move down a Group This trend occurs due to atoms more readily accepting electrons to fill their valence shells rather than losing them to have the below, already full, electron shell as their outer shell

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Notes

Predicting properties using the Periodic table Because there are patterns in the way the elements are arranged on the Periodic table, there are also patterns and trends in the chemical behaviour of the elements There are trends in properties down Groups and across a Period All of the Group I elements, for example, react very quickly with water

The Periodic Table | CIE IGCSE Chemistry Revision
Notes

The periodic table arranges all chemical elements in special ways. Different types of chemical bonding, and

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patterns and trends can be observed in their arrangement.

Patterns and trends in the periodic table - Periodicity ...
The Periodic Table. The Periodic table is a method of classifying elements. They are arranged in order of increasing atomic number. Made up of rows called periods and columns called groups. Metals are to the left and non-metals are to the right. Elements in the same group have similar properties.

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NOTES: 6.3 – Periodic Trends. PERIODIC TABLE:
RECALL... Periodicity:regular variations (or patterns)
of properties with increasing atomic weight; both
chemical and physical properties vary in a periodic
(repeating pattern). Group:vertical column of elements
(“ family ”) Period:horizontal row of elements.

CHAPTER 6 NOTES: The Periodic Table
Periodic Classification of Elements, Class 10 Chapter 5
Science Notes includes detailed explanations about the
modern periodic table, periodic trends and about its
discovery.

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In this lesson students learn about the trends of the Periodic Table by doing an inquiry activity, taking notes, watching videos, and doing practice questions. Within this unit students already learned about electronegativity in terms of bonding types in lesson 6: Metallic and Covalent Bonds.

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In the modern periodic table, the electronegativity of elements increases across a period (row) and decreases down a group (column). Therefore, the bottom-left most element (francium) is predicted to have the lowest electronegativity and the top-right

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most element (fluorine) is predicted to have the highest electronegativity.

Modern Periodic Table of Elements - Features ...

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